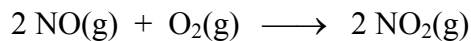


18. Consider the reaction:



The experimental Rate Law is:

$$\text{Rate} = k[\text{NO}][\text{O}_2]$$

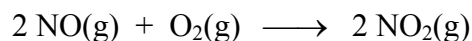
The order in NO is:

1. 2
2. 1
3. 1/2
4. 0

Correct Answer: 2.

Comments to the instructor: This concept tests has students distinguish between the rate law and the instantaneous rate.

19. Consider the reaction:



The experimental Rate Law is:

$$\text{Rate} = k[\text{NO}][\text{O}_2]$$

The overall order of the reaction is:

1. 2
2. 3
3. 1
4. 4

Correct Answer: 1.

Comments to the instructor: This concept test has students distinguish between the rate law and the instantaneous rate